**Chapter 2: The Chemical Context of Life**

**-AP Biology Guided Reading-**

***Sections 1 and 2***

1. Define and given an example of the following terms:
   1. Matter \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

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* 1. Element \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

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* 1. Compound \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

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1. What four elements make up 96% of all living matter?

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1. Explain the difference between an essential element and a trace element.

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1. What is the atomic number of Carbon? The atomic mass?

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1. Sketch a model of Carbon, showing the electrons, protons, neutrons, and atomic nucleus.
2. Provide the charges of each of the subatomic particles. Where is each particle located?
   1. Proton \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
   2. Neutron \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
   3. Electron \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
3. Define the following terms:
   1. Atomic mass \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

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* 1. Isotope \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

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1. Give a basic overview of how radioactive isotopes work.

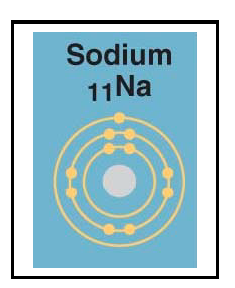
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1. What determines the chemical properties/behavior of an atom?

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1. Here is an electron distribution diagram for sodium:
   1. How many valence electrons does it have? Circle them.
   2. How many protons does it have?
   3. Assuming it is a neutral molecule, how many electrons does it have?
   4. Given that the mass number is 23, how many neutrons does it have?

***Sections 3 and 4***

1. Define molecule: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

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1. What does the term electronegativity mean?

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1. Explain the difference between a nonpolar covalent bond and a polar covalent bond.

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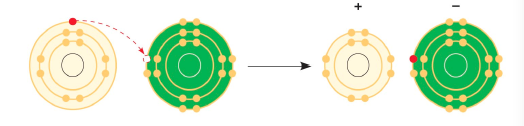
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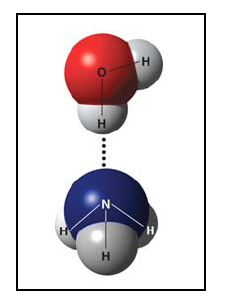
1. Does water form a nonpolar covalent bond, or a polar covalent bond? Explain why.

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1. Draw a diagram of a water molecule, labeling the regions that are more positive or more negative.
2. Another type of bond is shown here. State what type of bond this is, what two elements are involved, and explain what is happening in the diagram.
3. Explain the terms anion and cation, and how they are formed. In the previous example, which element is the anion, and which is the cation?

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1. What is a hydrogen bond? Indicate where the hydrogen bond occurs in this figure.
2. Describe van der Waals interactions.

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1. Place the following in order from weakest to strongest: van der Waals, covalent bonds, ionic bonds, hydrogen bonds.

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1. Why is molecular shape so important in cells?

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1. Write the chemical equation for photosynthesis, labeling the reactants and the products.
2. Explain dynamic chemical equilibrium.

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